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a. ammonium ions and sulfate ions b. sodium ions (group 1) and phosphate ions POH c. magnesium ions (group 2) and hydrogen sulfate ions d. aluminum ions (group 13) and carbonate ions COD 3 onium ions and arsenate ions (AsOq0 f. calcium ions (group 2) and acetate ions 71 g. ammonium ions and nitrite ions Math Skills Transparency Worksheets

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ion is then free to move throughout the molten mass. If a voltage is applied, cations will migrate to one electrode, and anions will migrate to the other. This movement of ions means that there is a flow of electricity between the two electrodes. When ionic compounds dissolve in water, their ions are free to move. Thus, aqueous solutions of ionic

Part D Questions and Problems 24. a. b. 25. a. b. c. d. 2. ...
In general, ionic compounds have high melting points because the ions have a strong attraction for one another. Ionic compounds conduct an electric current when melted or in an aqueous solution because the ions are then free to move. After reading Lesson 7.2, answer the following questions. Formation of Ionic Compounds 1.

BONDING AND INTERACTIONS
As you have learned, ions are atoms or molecules bearing an electrical charge. A cation (a positive ion) forms when a neutral atom loses one or more electrons from its valence shell, and an anion (a negative ion) forms when a neutral atom gains one or more electrons in its valence shell. ... Answers to Chemistry End of Chapter Exercises. 1. The ...

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The ions attract, forming Al 25 3. 2 Al ions (_ 3 Al ions) 3 S ions (_ 2 S ion) 2(3) 3(2) 0 The overall charge on one formula unit of Al 25 3 is zero. 11. Challenge Explain how elements in the two groups shown on the periodic table at the right combine to form an ionic compound. Group 1 Group 15 Three group 1 atoms loses 1 e, forming 1 ions.

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